Classifying Living Things

Over 2000 years ago, the Greek philosopher Aristotle developed a system of classification that grouped organisms according to whether they were plant or animal. Scientists used Aristotle's system for hundreds of years, but as they discovered more and more living things, the system did not work well because it did not show probable relationships between similar organisms.

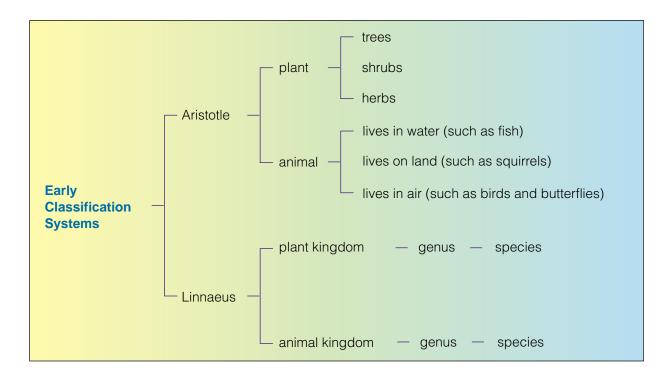
In 1735, Carolus Linnaeus produced a new system that also classified all organisms as plant or animal, but this new system was very different in other ways from Aristotle's system.

Linnaeus' system gives a two-word name to each type of organism. This system of naming organisms is still in use today. The two-word name is called the organism's scientific name,

organisms is still in use today. The two-word name is called the organism's scientific name, and it is given in Latin, a language that is no longer spoken. The first word of the organism's name is its genus, and the second word is its specific name. A **genus** is a group of organisms that are very similar. A **species** is the smaller, more limiting classification grouping. The Canada lynx, shown on the right, and the bobcat, shown on the left, are members of the same genus, *Lynx*, but they are considered different species. The lynx is the species *Lynx canadensis* and the bobcat is *Lynx rufus*.





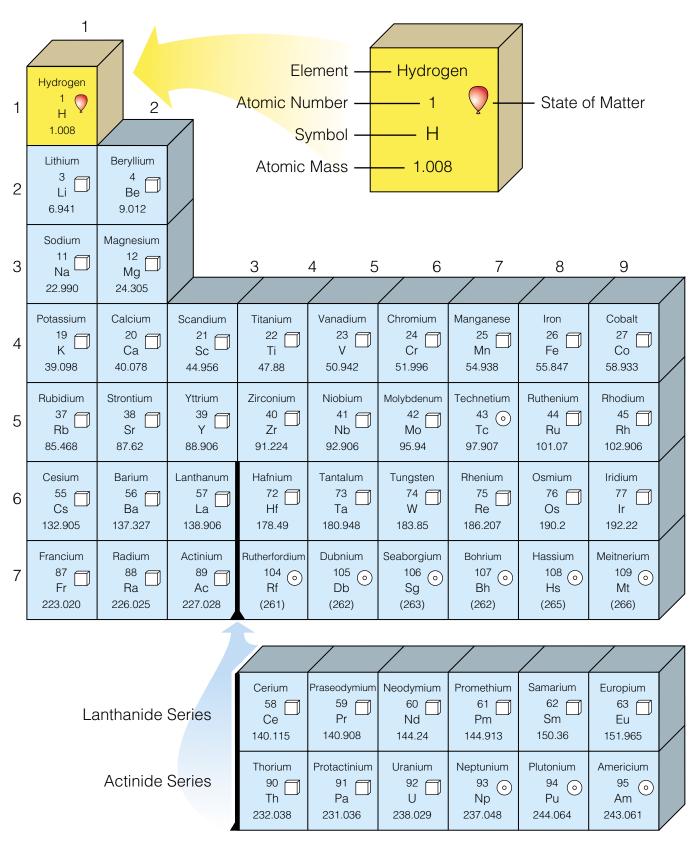


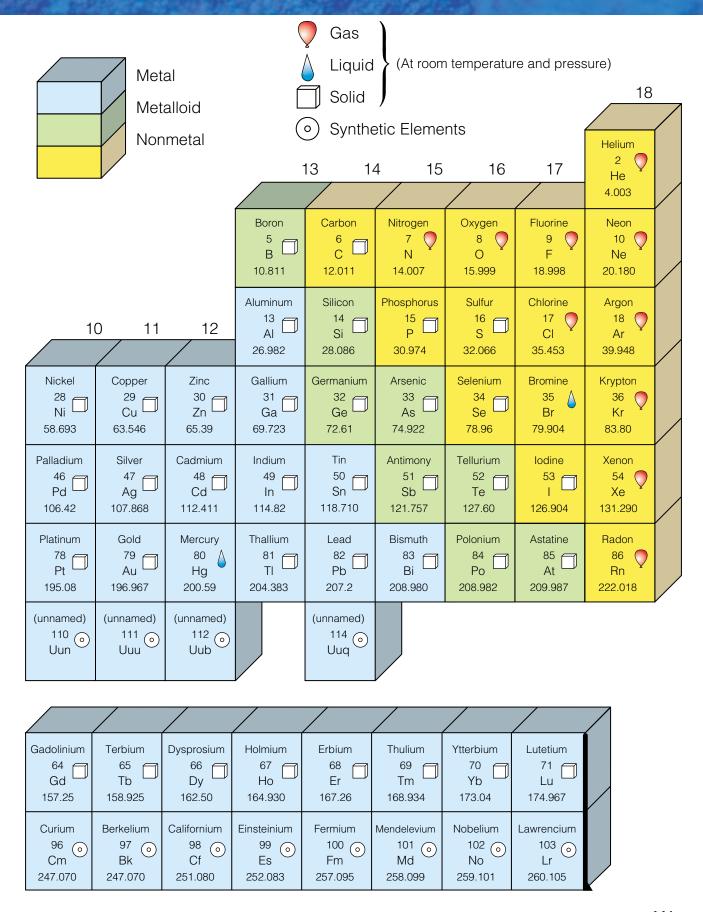
In the 1900s, as knowledge about the great diversity of organisms on Earth exploded, it became clear that separating organisms into only two kingdoms, plant and animal, was inadequate. For example, bacteria are just too different from either plants or animals to be grouped with either. Likewise, fungi such as bread mould, yeast, and the many kinds of mushrooms are very different from plants and animals. Influential scientists like Robert Whittaker and Lynn Margulis supported the idea that new criteria for classifying organisms were needed. Thus, since the 1960s, a system that classifies organisms into five different kingdoms, still using Linnaeus' basic system at its roots, is largely accepted and used. The organisms and their kingdoms are shown in the table below.

Life's Five Kingdoms

| | Monera | Protista | Fungi | Plant | Animal |
|-----------------|--|--|---|------------------------------|--|
| Number of cells | One-celled | One- and many-celled | One- and many-celled | Many-celled | Many-celled |
| Movement | Some move | Some move | Don't move | Don't move | Move |
| Nutrition | Some members make their own food; others obtain it from other organisms. | Some members make their own food; others obtain it from other organisms. | All members obtain food from other organisms. | Members make their own food. | Members eat plants or other organisms. |
| | | | | | |

PERIODIC TABLE OF THE ELEMENTS





Properties of Common Substances

KEY TO SYMBOLS:

Common names of substances are enclosed in parentheses.

(*) Water solution of a pure substance

(e) Element

(c) Compound

(n) Not a pure substance

| () | (-) | (-) (-) | () | | |
|-------------------------------------|---|-----------------------|---|--|--|
| Name | Formula | Melting point (°C) | Boiling point (°C) | Density (g/cm³ or g/mL) | |
| acetic acid (vinegar) (c) | CH ₃ COOH | 16.6 | 118.1 | _ | |
| alcohol (see ethanol) (c) | Ŭ | | | | |
| aluminum (e) | Al | 659.7 | 2519 | 2.7 | |
| aluminum oxide (alumina) (c) | Al ₂ O ₃ | 2015 | _ | _ | |
| ammonia (c) | NH ₃ | -77.8 | -33.4 | less dense than air | |
| ammonium nitrate (c) | NH_4NO_3 | 169.6 | 210 | | |
| antimony (e) | Sb | 631 | 1587 | 6.70 | |
| argon (e) | Ar | -189 | -185 | denser than air | |
| arsenic (e) | As | _ | _ | 5.727 (grey) 4.25 (black) 2.0 (yellow) | |
| barium (e) | Ва | 727 | 1897 | 3.62 | |
| berkelium (e) | Bk | 1050 | _ | 14.78 | |
| beryllium (e) | Be | 1280 | 2471 | 1.85 | |
| bismuth (e) | Bi | 271 | 1560 | 9.7 | |
| boron (e) | В | 2075 | 4000 | 2.37(brown) 2.34 (yellow) | |
| bromine (e) | Br | -7.2 | 58.8 | 3.12 | |
| calcium (e) | Ca | 845 | 1484 | 1.55 | |
| calcium carbonate (limestone) (c) | CaCO ₃ | decomposes at 900°C | _ | 2.93 | |
| calcium hydroxide (slaked lime) (c) | Ca(OH) ₂ | decomposes at 522°C | _ | 2.24 | |
| calcium oxide (lime) (c) | CaO | 2580 | 2850 | 3.3 | |
| carbon (diamond) (e) | С | 3500 | 4827 | 3.51 | |
| carbon (graphite) (e) | С | 3652–3697 | 4827 | 2.25 | |
| carbon dioxide (c) | CO ₂ | _ | _ | _ | |
| chlorine (e) | Cl ₂ | -101.6 | -34.6 | denser than air | |
| chromium (e) | Cr | 1907 | 2671 | 7.2 | |
| cobalt (e) | Со | 1480 | 2927 | 8.9 | |
| copper (e) | Cu | 1084 | 2562 | 8.95 | |
| copper (II) nitrate (c) | $Cu(NO_3)_2$ | _ | _ | _ | |
| copper (II) sulfate (bluestone) (c) | CuSO ₄ .5H ₂ O | decomposes at 150°C | _ | 2.28 | |
| diamond (see carbon) (e) | 7 - | | | | |
| ethanol (ethyl alcohol) (c) | C₂H₅OH | -114.5 | 78.4 | 0.789 | |
| ethylene (ethene) (c) | C_2H_4 | -169 | -103.9 | _ | |
| fluorine (e) | F ₂ | -270 | -188 | _ | |
| gold (e) | Au | 1063 | 2856 | 19.3 | |
| glucose (c) | C ₆ H ₁₂ O ₆ | 146 | decomposes before it boils | 1.54 | |
| graphite (see carbon) (e) | 0 12 0 | | , | | |
| helium (e) | He | -272.2 | -268.93 | _ | |
| hematite (c) | Fe ₂ O ₃ | 1565 | | 5.24 | |

DEFINITIONS:

 $\label{eq:continuous} \begin{tabular}{ll} deliquescent — able to absorb water from the air to form a concentrated solution sublime — to form a vapour directly from a solid \\ \end{tabular}$

| (at room temperature: 20°C) colourless liquid with pungent smell used in the manufacture of cellulose ethanoate; vinegar is a 5–7 percent solution is silver-white metal white, crystalline substance used in aircraft, cooking utensils, and electrical apparatus used in refining of aluminum and in cement very soluble gas with pungent smell used as refrigerant and in manufacture of resins, explosives, and fertilizers white, soluble, crystalline salt used in explosives and as a fertilizer silver-grey solid used in infra-red detectors inert gas used in electric lights grey, black, or yellow solid used in semiconductors and alloys; compounds are very poisonous and are used in semiconductors and alloys; compounds are very poisonous and are used in semiconductors. | |
|---|----------|
| silver-white metal used in aircraft, cooking utensils, and electrical apparatus white, crystalline substance used in refining of aluminum and in cement very soluble gas with pungent smell used as refrigerant and in manufacture of resins, explosives, and fertilizers white, soluble, crystalline salt used in explosives and as a fertilizer silver-grey solid used in infra-red detectors inert gas used in electric lights | |
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| silver-grey solid used in infra-red detectors inert gas used in electric lights | ed in |
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| | ed in |
| medicine and as pesticides | |
| silver-white solid used in X-ray diagnosis | |
| | |
| hard, white metal used for corrosion-resistant alloys | |
| brittle, white, crystalline metal with reddish tinge used in alloys, catalysts, nuclear reactors; compounds used in medicine | |
| brown amphorous powder or yellow crystals used for hardening steel and for producing enamels and glasses | |
| red-brown liquid liquid causes severe chemical burns; vapour is harmful to lungs; used to make pain-relieving drugs | certain |
| soft, white metal that tarnishes easily very abundant; essential to life | |
| white solid main ingredient in chalk, marble | |
| white solid $ {\it aqueous solution used to test for CO}_2 $ | |
| white solid used in cement and to mark lines on playing fields | |
| colourless, solid crystals very hard; used for drilling through rock | |
| grey-black solid very soft; used in lubricants, pencil leads, electrical apparatus | |
| colourless gas with a faint tingling smell and taste does not support combustion and is denser than air; used in fire extinguishers as a refrigerant at -78.5°C | and |
| green gas poisonous; used to kill harmful organisms in water | |
| shiny, silvery solid very hard metal; used to make stainless steel | |
| hard, silver-white, magnetic metal used in alloys; compounds used to produce the colour blue in glass and ceram | ics |
| shiny, reddish solid soft metal; good conductor of heat | |
| | |
| blue, solid crystals used in pesticides | |
| | |
| colourless liquid derived from fermentation of sugar; used as solvent or fuel; found in wine | i a a la |
| colourless, flammable gas with a sweetish smell made from petroleum; used in manufacture of ethanol and other organic chem | icais |
| greenish yellow gas similar to chlorine | |
| shiny, yellow solid very soft metal; highly resistant to tarnishing | |
| white solid simple sugar; human body converts most sugars and starches to glucose | |
| poplommobile input gos | hallassa |
| nonflammable inert gas used as refrigerant; provides inert atmosphere for welding; used to fill air ships and | Dalloons |
| rusty red colour found in iron ore; rusty iron | |

| | | Melting point | Boiling point | Density | |
|-----------------------------------|---|---------------|---------------------|------------------------------|--|
| Name | Formula | (°C) | (°C) | (g/cm³ or g/mL) | |
| hydrochloric acid (*) | HCI | varies | varies | varies | |
| hydrogen (e) | H ₂ | -259 | -253 | much less dense than air | |
| hydrogen peroxide (c) | H ₂ O ₂ | -0.4 | 150.2 | 1.45 | |
| iodine (e) | 11202 | 114 | 184 | 4.95 | |
| iron (e) | Fe | 1535 | 2861 | 7.86 | |
| lead (e) | Pb | 327.4 | 1750 | 11.34 | |
| lead (II) nitrate (c) | Pb(NO ₃) ₂ | 327.4 | — — — | 11.04 | |
| limestone (see calcium carbonate) | FD(NO ₃) ₂ | _ | _ | _ | |
| lithium (e) | Li | 179 | 1340 | 0.534 | |
| magnesium (e) | Mg | 651 | 1107 | 1.74 | |
| magnesium (e) | MgCl ₂ | 708 | 1412 | 2.3 | |
| magnetite (c) | | — — | — 1412 — | 5.18 | |
| | Fe ₃ O ₄ | 1246 | 2061 | 7.43 | |
| manganese (e) | | -38.9 | 356.6 | 13.6 | |
| mercury (e) | Hg | | | | |
| methane (c) | CH ₄ | -182.5 | -161.5 | _ | |
| molybdenum (e) | Мо | 2623 | 4679 | 10.28 | |
| neon (e) | Ne | -248 | -246 | _ | |
| nickel (e) | Ni | 1455 | 2913 | 8.90 | |
| nitrogen (e) | N_2 | -209.9 | -195.8 | slightly less dense than air | |
| nitrogen dioxide (c) | NO ₂ | _ | _ | _ | |
| oxygen (e) | 02 | -218 | -183 | slightly denser than air | |
| ozone (e) | 03 | -192.5 | -112 | denser than air | |
| phosphorus (e) | Р | 44 | 280 | 1.82 (white) | |
| p | | _ | | 2.20 (red) | |
| | | | | | |
| platinum (e) | Pt | 1769 | 3824 | 21.41 | |
| polyethylene (polythene) (c) | $(C_2H_4)_n$ | _ | _ | _ | |
| potassium (e) | K | 63.5 | 759 | 0.86 | |
| propane (c) | C ₃ H ₈ | _ | -42.17 | _ | |
| selenium (e) | Se | 217 | 684.9 | 4.81 | |
| silicon (e) | Si | 1410 | 3265 | 2.33 | |
| silicon dioxide (silica) (c) | SiO ₂ | 1600 | _ | _ | |
| silver (e) | Ag | 961 | 2212 | 10.5 | |
| sodium (e) | Na | 97.5 | 892 | 0.971 | |
| sodium chloride (table salt) (c) | NaCl | 801 | 1465 | 2.16 | |
| sodium fluoride (c) | NaF | 988 | 1695 | 2.56 | |
| steel (n) | varies | varies | varies | varies | |
| strontium (e) | Sr | 777 | 1412 | 2.6 | |
| sucrose (sugar) (c) | C ₁₂ H ₂₂ O ₁₁ | 170 | decomposes at 186°C | 1.59 | |
| sulfur (brimstone) (e) | S S | 112.8 | 444.6 | 2.07 | |
| technetium (e) | Tc | 2157 | 4265 | 11.5 | |
| tellurium (e) | Te | 450 | 990 | 6.25 | |
| tin (e) | Sn | 231.9 | 2602 | 7.31 | |
| titanium (e) | Ti | 1666 | 3287 | 4.5 | |
| tungsten (e) | W | 3422 | 5555 | 19.25 | |
| uranium (e) | U | 1130 | 4131 | 19.05 | |
| water (c) | H ₂ 0 | 0 | 100 | 1.00 | |
| xenon (e) | Xe | -111.9 | -107.1 | _ | |
| zinc (e) | Zn | 419 | 907 | 7.14 | |
| zirconium (e) | Zr | 1852 | 4400 | 6.51 | |
| (-) | | .00= | | 0.0 | |

| colourless liquid colourless liquid colourless liquid colourless liquid colourless liquid description in the colourless liquid sole-black, sella Crystals shiny, silver solid shiny, blue-white solid shiny, blue-white solid sole-black, sella (crystals sole-black) sole-black, sella (crystals solid sole-black) sole-black, sella (crystals sole-black, sella (crystals sole-black) sole-black, sella (crystals) sole-black, sella (cry | Appearance (at room temperature: 20°C) | Comments |
|--|--|---|
| colouriess gas diphty fammable: floud form is used as nocket fuel disches floud shiny, silver vehitle sold shiny, silver vehitle sold shiny, silver vehitle sold sold metal. forms poisonous compounds whito or colouriess crystals saliver-whitle metal (least dense solid known) light, silvery white metal dispersion of the solid grey-white solid grey-white solid grey-white solid grey-white solid grey-white solid grey-white solid gray in the so | | |
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| silver-white metal (least dense solid known) light, silvery white metal (least dense solid known) light, silvery white metal (least dense solid known) white, deliquescent substance shiny, back crystaline solid grey-white solid shiny, silvery liquid only liquid metal forms poisonous compounds shiny, silvery liquid on liquid metal forms poisonous compounds odoruress, flammable gas formed from decaying organic matter silver-white solid colouriess, goanness gas silver-white solid colouriess, goanness gas silver-white solid colouriess, goanness gas silver-white, anginetic metal that resists corrosion colouriess gas causes redish brown gas colouriess gas colour | • | |
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